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stoichiometry

problem, the

first step is

always to

calculate the

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number of moles  
of each reactant  
present. In this  
case, we are  
given the mass  
of  $K_2Cr_2O_7$   
in 1 mL of  
solution, which  
we can use to  
calculate the  
number of moles  
of  $K_2Cr_2O_7$  ...

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review. What is the mass in grams of  $15 \times 10^{16}$  atoms s.

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What are the reactants.

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Mole Ratios

Questions 1.

Aluminum reacts with oxygen to produce aluminum oxide as

follows:  $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$

a. If you use 2.3 moles of Al, how many moles

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of Al<sub>2</sub>O<sub>3</sub> can you make? b. If you want 3.9 moles of Al<sub>2</sub>O<sub>3</sub>, how many moles of O<sub>2</sub> are needed? 2. In the presence of sulfuric acid, metallic iron forms iron(III) sulfate:  $2\text{Fe} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2$  ...

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mass. What does  
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